Group Members:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date:\_\_\_\_\_\_\_\_\_\_\_\_

In Groups of Four, work as a team to balance the following redox equations:

Balance the following chemical equations for reactions in an acidic solution.

1. IO3-(aq) + HSO3-(aq) → SO42-(aq) + I2(aq)

Balance the following chemical equations for reactions on a basic solution.

1. ClO3-(aq) + N2H4(aq) → NO(g) + Cl-(aq)

TEACHER COPY

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Balance the following chemical equations for reactions in an acidic solution.

1. IO3-(aq) + HSO3-(aq) → SO42-(aq) + I2(aq)

+5-2 +1+4-2 +6 -2 0

IO3-(aq) + HSO3-(aq) → SO42-(aq) + I2(aq)

2 IO3-(aq) + 5 HSO3-(aq) → 5 SO42-(aq) + I2(aq)

2 IO3-(aq) + 5 HSO3-(aq) → 5 SO42-(aq) + I2(aq) + H2O(l)

2 IO3-(aq) + 5 HSO3-(aq) → 5 SO42-(aq) + I2(aq) + H2O(l) + 3H+

Balance the following chemical equations for reactions on a basic solution.

1. ClO3-(aq) + N2H4(aq) → NO(g) + Cl-(aq)

+5 -2 -2 +1 +2 -2 -1

ClO3-(aq) + N2H4(aq) → NO(g) + Cl-(aq)

4 ClO3-(aq) + 3 N2H4(aq) → 6 NO(g) + 4 Cl-(aq)

4 ClO3-(aq) + 3 N2H4(aq) → 6 NO(g) + 4 Cl-(aq) + 6H2O(l)